Periodic Table Trends



Atomic Radius

* Increases at it moves down a group
* Decreases as it moves from the left to the right in a period

Ionization Energy

* Increases as it moves up a group and from left to right in a period

Electron Affinity

* Increases as it moves up a group and from left to right in a period

Electronegativity

* Increases as it moves up a group and from left to right in a period
* Noble gases have zero electronegativity

Atomic Mass

* Increases from left to right on the periods and from top to bottom in a group

Reactivity

* Metals and non-metals have separate trends. Metal’s reactivity increases as you move down a group and generally decreases as you move left to right across a period.
* Non-metals are less reactive as you move down a group and reactivity generally increases moving left to right.

 

*Note:*

\*Melting points and Density have no set trends