**PeRiodIc TaBlE tErMs!!!**

**Atomic Radius**

* Half the distance between the nuclei of two bonded atoms (this gives an idea of the size of the atom).
* Atomic radius is influenced by the strength of attraction between the nucleus and the electrons in the outermost energy level. If the attraction is strong, the atom is smaller; if the attraction is weak, the atom is larger.

**Ionization Energy**

* The energy required to remove an electron from an atom or ion.
* The closer an electron is to the nucleus, the more difficult it is to remove, and the higher its ionization energy will be.
* 1st ionization energy is the energy required to remove an electron from a neutral atom while 2nd ionization energy is the energy required to remove an electron from an ion.

**Electron Affinity**

* A measure of the energy given off when a gaseous atom gains an electron and forms a negatively charged ion.
* The greater an atom’s electron affinity, the easier it is for the atom to accept an electron.
* If an atom gains an electron and becomes more stable => releases energy & electron affinity value is negative.
* If an atom gains an electron and becomes less stable => absorbes energy and affinity value is positive

**Electronegativity**

* Tendency of an atom to attract electrons in a chemical bond.
* Measured on the Pauling scale and ranges from francium (0.7) to fluorine (4.0)

**Atomic Mass**

* The sum of the protons and neutrons in the nucleus of an atom.

**Melting Point**

* The temperature at which a solid becomes a liquid. Most elements are solid at room temperature and have a high melting point. Elements that are gaseous at room temperature have a melting point below 0°C.

**Reactivity**

* Refers to how vigorously an atom is to react with other substances.
* Usually determined by how easily electrons can be removed (ionization energy) and how badly they want to take other atom’s electrons (electronegativity).